Total No. of Questions – [4]

Total No. of Printed Pages: – [2]

P118-122(T1)

G.R. No.	

OCTOBER 2018 / IN - SEM (T1)

F. Y. M. TECH. (WREE) (SEMESTER -I)

COURSE NAME: Environmental Chemistry and Microbiology

COURSE CODE: CVPA11182

(PATTERN 2018)

Time: [1 Hour] [Max. Marks: 20]

(*) Instructions to candidates:

1) **Answer Q.1 OR Q.2, Q.3 OR Q.4**
2) **Figures to the right indicate full marks.**
3) **Use of scientific calculator is allowed**
4) **Use suitable data where ever required**

Q.1)a)The dissociation constant of weak base is 3.2×10^{-2}.Determine degree of dissociation when concentration is 0.05 Molar [4]

b) Explain terms used in Dulong's formula and determine calorific value of atomic structure $C_6H_8N_3S_{20}$. [6]

OR

Q.2)a) What is concentration of chlorine gas in water at 20 ^{0}C temperature. Take partial pressure of chlorine gas 0.2 atm and Henry's Constant H $=4 \times 10^{-5}$ [4]

b)Write short note on Alkenes and Alkanes. [6]

Q.3) a)What is current in amperes required to liberate 2×10^{-3}Kg of iodine from potassium iodide in 2 hours. Take Eq. Mass of Iodine =130 [4]